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**Process for the catalyst-free preparation of
cyanophenols from methoxybenzonitriles**

Description

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The present invention provides a process for the catalyst-free preparation of cyanophenols.

10 A common, industrially established process for preparing phenols which is described in detail in the literature [K. Weissermel, H.-J. Arpe, Industrielle organische Chemie [Industrial organic chemistry], 3rd revised and extended edition, VCH Weinheim, 1988, p. 376] is the known conversion of an isopropylaromatic
15 using oxygen and the subsequent conversion to the phenol and acetone (Hock phenol synthesis). Disadvantages of this process are the equimolar occurrence of acetone and the high consumption of propene to prepare the isopropylaromatic. In addition,
20 it is not possible to prepare arbitrarily substituted cyanophenols by this process.

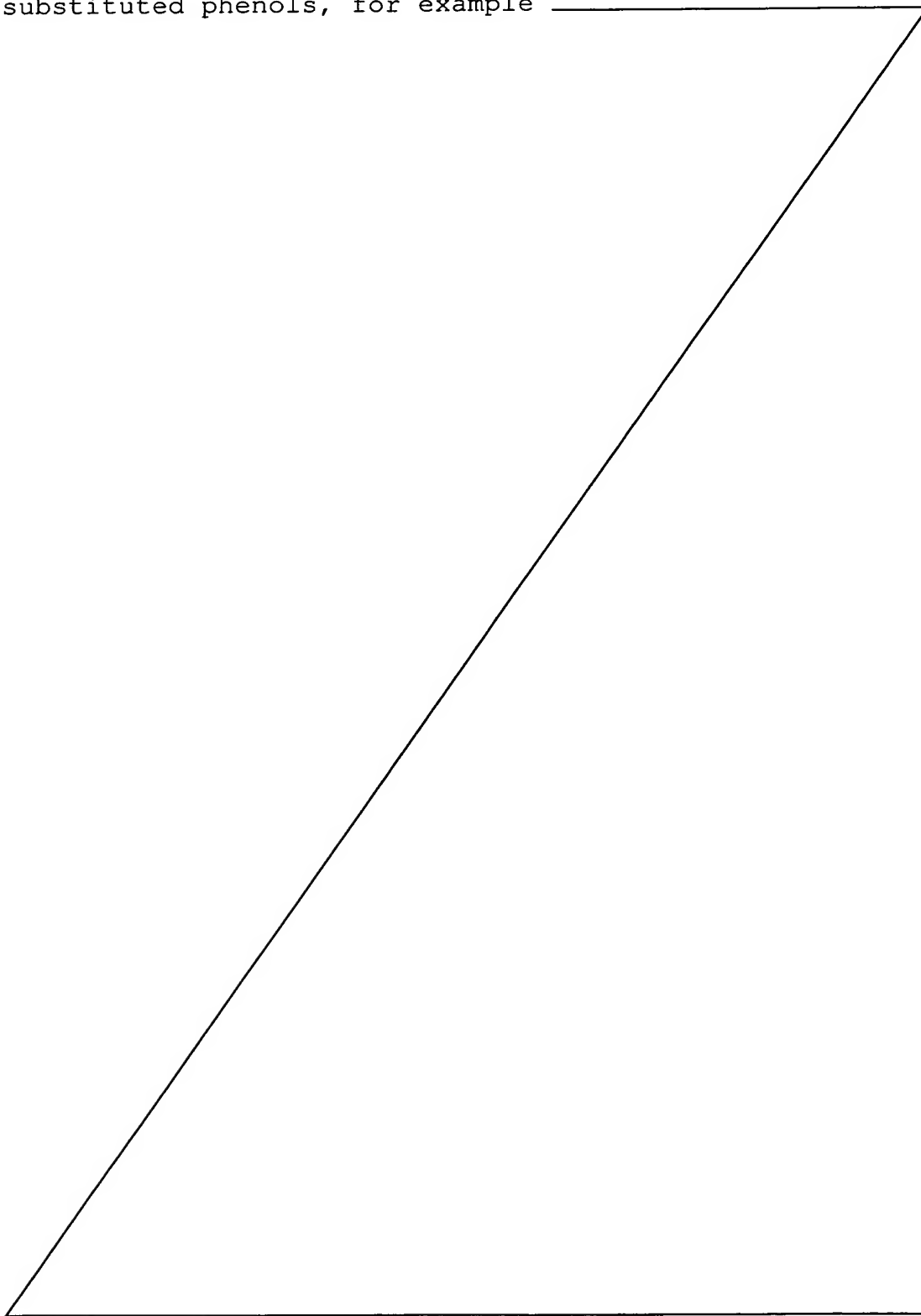
In the already classic process [K. Weissermel, H.-J. Arpe, Industrielle organische Chemie [Industrial
25 organic chemistry], 3rd revised and extended edition, VCH Weinheim, 1988, p. 370] for preparing phenols from benzenesulfonic acids, large amounts of salts (Na_2SO_3 , Na_2SO_4) are obtained as a coproduct, so that the process can no longer be used in the western world for
30 environmental reasons.

The preparation of phenols from chloroaromatics by means of NaOH is also known from the literature and has been industrially realized [K. Weissermel, H.-J. Arpe,
35 Industrielle organische Chemie [Industrial organic chemistry], 3rd revised and extended edition, VCH Weinheim, 1988, p. 372], but in particular the

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necessary, sometimes drastic reaction conditions and particularly the temperatures above 360°C are regarded as disadvantageous. A conversion of the process to substituted phenols, for example

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Also known from the literature is the ammoxidation of methylphenols to the corresponding cyanophenols [M. V. Landau et al., Applied Catalysis A: General (2001), 208(1,2), 21-34; A. Martin et al., J. Prakt. Chem. (1990), 332(4), 551-6; H. Bruins Slot, DE 20 37 945], which succeeds, though, in only very modest yields. Disadvantages of this process are not only the low yield but also the high cost of inconvenience for the industrial synthesis. The same applies for the formation of the nitrile function from a benzoic acid and ammonia or from an ester and ammonia, according to the following references [R. Ueno et al., EP 74 116; M. Araki et al., JP 53040737; R. Perron, FR 2 332 261; G. Bakassian, M. Lefort, DE 22 05 360; H. Eilingsfeld, E. Schaffner, DE 20 20 866; T. Ichii et al., JP 43029944].

The literature also gives details on the ether cleavage to form phenols [P. R. Brooks et al., Journal of Organic Chemistry (1999), 64(26), 9719-9721]. In this process, especially the stoichiometric use of the expensive raw materials boron trichloride and the use of large amounts of n-butylammonium iodide are regarded as disadvantageous.

Systems which have been developed specifically for the cleavage of allyl ethers are also described in the present context. To this end, it is possible to use systems composed of CeCl_3 and NaI [R. M. Thomas et al., Tetrahedron Letters (1999), 40(40), 7293-7294] or NaBH_4 [R. M. Thomas et al., Tetrahedron Lett. (1997), 38(26), 4721-4724], or else electrochemical processes

[D. Franco et al., Tetrahedron Lett. (1999), 40(31), 5685-5688; A. Yasuhara et al., J. Org. Chem. (1999), 64(11), 4211-4213; K. Fujimoto et al., Tetrahedron (1996), 52(11), 3889-96; S. Olivero et al., J. Chem. Soc., Chem. Commun. (1995), (24), 2497-8]. However, the latter process variant is expensive and it always requires the use of heavy metals; in addition, this process variant is restricted exclusively to the allyl ethers which are difficult to synthesize.

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The preparation of nitrophenols from nitroaromatics by substituting hydrogen with hydroperoxide anions in the presence of strong bases is a very interesting process, but is unfortunately restricted to nitroaromatics; in addition, it is necessary to use liquid ammonia and readily decomposable and thus dangerous hydroperoxides [M. Makosza et al., J. Org. Chem. 1998, 63, 4199-4208]. To prepare cyanophenols, the nitro group has to be converted to a cyano group in a complicated manner.

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T. Senba and K. Sakano (JP 09023893) and H. Semba et al. describe the enzymatic synthesis of phenols [Appl. Microbiol. Biotechnol. 1996, 46, 432-437]. However, the low space-time yields and the long reaction times mean that this process cannot be utilized in an economic manner.

The synthesis of phenols from anilines by diazotization and decomposition of the diazonium compound in the presence of metals, particularly copper salts, has also been known for some time. Even though more recent studies have attempted to optimize the process [B. C. Gilbert et al., EP 596 684], this route nevertheless always leads via a diazonium compound which is difficult to handle.

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According to S. Prouilhac-Cros et al., the preparation of phenols from arylsilanes with H_2O_2 and stoichiometric

amounts of fluoride leads to phenols in good yields [Bull. Soc. Chim. Fr. 1995, 132, 513-16]. However, the arylsilane required is not sufficiently available in industrial quantities.

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The synthesis of phenols from aryl methyl ethers with ethanethiolates is restricted to laboratory applications [J. A. Dodge et al., J. Org. Chem. 1995, 60, 739-41]. For industrial syntheses, the reaction is unsuitable because toxic and malodorous sulfur compounds are obtained.

The cleavage of aryl methyl ethers with the FeO/glacial acetic acid/oxygen system has likewise been described [A. F. Duprat et al., J. Mol. Catal. 1992, 77]. This process is restricted to particularly activated aromatics and gives the desired products only in low to very low yields. Only slightly better yields in the cleavage of aryl methyl ethers are obtained with AlCl₃/NaCl [G. Adamska et al., Biul. Wojsk, Akad. Tech. 1980, 29, 93-99]. The large amounts of inorganic waste which form do not allow any industrial application of the process. The AlCl₃/Ni system which is described in the literature for the cleavage of these ethers also does not offer any advantages whatsoever, since temperatures above 240°C and long reaction times are required [H. Kashiwagi, S. Enomoto, Yakugaku Zasshi 1980, 100, 668-71].

A very complicated process for the cleavage of aryl methyl ethers and for the preparation of hydroxybenzonitrile is transmethylation (for example FR 1 565 812), in which the reaction of methoxybenzonitrile with the sodium salt of cresol to the sodium salt of hydroxybenzonitrile and methoxycresol is effected at temperatures above 200°C. This process affords large amounts of waste and has a very complicated procedure. Although aryl methyl ethers in principle constitute

very suitable raw materials, the cleavage of the ether is very difficult; alternative cleavage processes are not available.

5 US-A-3,567,758 describes a process for preparing 4-alkoxybenzonitriles, in which, however, alkoxides are not used, but rather sodium 4-methylphenoxide (sodium phenoxide).

10 Alkoxides have in fact already been used for the synthesis of phenols (see, for example, Testaferri et al., The Reaction of Unactivated Aryl Halides with Sodium Methoxide in HMPA, Tetrahedron 39, 1, 193-197 (1983)). However, this publication does not address in
15 any depth the preparation of cyanophenols.

GB 1,065,936 from 1967 discloses a process for preparing 4-hydroxybenzonitrile by reacting anisic acid with a nitrile.

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EP 0 421 385 discloses a process for preparing an hydroxyneophyl m-phenoxybenzyl ether using a lower alkoxide. Here too, however, there is no information on the preparation of cyanophenols.

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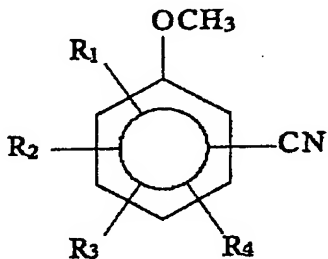
All processes mentioned lead to the desired products and have also already been used successfully for a large number of highly varying applications. However, all of these processes for preparing phenols and
30 especially cyanophenol have the disadvantage that they have a very complicated procedure, that expensive raw materials have to be used, that large amounts of waste are formed or that the yield is only very low.

35 It is therefore an object of the present invention to develop a process for the catalyst-free preparation of cyanophenols, which allows the environmentally friendly

and low-waste preparation of the desired product with favorable raw materials in high yields. In particular, the use of heavy metals, as is customary, for example, when metals are used as the catalyst, should be
5 dispensed with.

This object is achieved by a corresponding process, in which a substituted methoxybenzonitrile of the general formula (I)

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where

15 R_1 , R_2 , R_3 and R_4 are each independently hydrogen, a C_{1-10} -alkyl, C_{2-8} -alkoxy, aryl, a phenoxy or a further nitrile group
is reacted with an alkali metal alkoxide at temperatures between 80 and 230°C.

20 It has been found, surprisingly, that not only is it possible, as desired, to dispense fully with catalysts for the performance of the reaction, and that the cyanophenols are obtained in very good yields and purities, but also that a relatively simple process
25 which can be performed without the occurrence of by-products on the industrial scale without any problems is thus available.

30 The selection of the aromatic raw materials is not restricted only to simple methoxybenzonitriles, but rather also includes substituted methoxybenzonitriles, and in particular di-, tri-, tetra- or pentamethoxy-

benzonitriles are suitable.

Preferred alkali metal alkoxide components are methoxides and, among these, in particular sodium
5 methoxide.

The process according to the invention can be carried out within a relatively large temperature range. However, particularly suitable reaction temperatures
10 have been found to be between 120 and 200°C and most preferably between 140 and 180°C.

Typically, the reaction succeeds at best when the molar ratio of the methoxybenzonitrile component to the
15 alkali metal alkoxide component is from 1:0.5 to 1.5 and more preferably 1:1.0 to 1.1.

Typically, the present process succeeds even without the presence of a solvent. However, the present
20 invention also provides for the use of a suitable solvent, for which either polar or nonpolar solvents may be used. Particular preference is given to C₁₋₆-alcohols, for example methanol, and/or a solvent from the group of tetrahydrofuran, benzene, toluene, xylene
25 and methyl tert-butyl ether. Especially suitable for practical reasons and for reasons of cost is the use of simple alcohols, for example methanol.

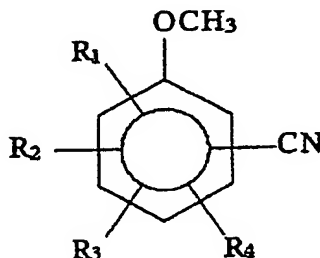
The preferred reaction is typically carried out by
30 initially charging the alkoxide component in an alcohol (for example sodium methoxide in methanol), then adding the methoxybenzonitrile component, which is preferably effected with stirring; this should preferably be effected in an autoclave, in which case

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What is claimed is:

1. A process for the catalyst-free preparation of cyanophenols from methoxybenzonitriles, characterized in that a substituted methoxybenzonitrile of the general formula (I)



- where
R₁, R₂, R₃ and R₄ are each independently hydrogen, a C₁₋₁₀-alkyl, C₂₋₈-alkoxy, aryl, a phenoxy or a further nitrile group
- is reacted with an alkali metal alkoxide at temperatures between 80 and 230°C.
2. The process as claimed in claim 1, characterized in that the methoxybenzonitrile component used comprises di-, tri- tetra- or pentamethoxybenzonitriles.
3. The process as claimed in either of claims 1 and 2, characterized in that the alkali metal alkoxide used is a methoxide, more preferably sodium methoxide.
4. The process as claimed in one of claims 1 to 3, characterized in that it is carried out at temperatures between 120 and 200°C and more preferably between 140 and 180°C.

5. The process as claimed in one of claims 1 to 4,
characterized in that the molar ratio of the
methoxybenzonitrile component to the alkali metal
alkoxide component is 1:0.5 to 1.5 and more
preferably 1:1.0 to 1.1.
6. The process as claimed in one of claims 1 to 5,
characterized in that it is carried out in the
presence of a polar and/or nonpolar solvent, more
preferably in the presence of a C₁₋₆-alcohol, e.g.
methanol, and/or of a solvent from the group of
tetrahydrofuran, benzene, toluene, xylene and
methyl tert-butyl ether.
7. The process as claimed in one of claims 1 to 6,
characterized in that the alkoxide component is
initially charged in an alcohol, then the methoxy-
benzonitrile component is added and preferably
heated with stirring, which is more preferably
effected in an autoclave.
8. The process as claimed in one of claims 1 to 7,
characterized in that the methoxybenzonitrile
component has been prepared by ammoxidizing a
methoxytoluene and in the presence of ammonia and
(atmospheric) oxygen.
9. The process as claimed in claim 8, characterized
in that the methoxybenzonitrile component is
reacted further after the ammoxidation directly
without isolation.